UNIVERSITAS NEGERI YOGYAKARTA
FACULTY OF MATHEMATICS AND NATURAL SCIENCES
DEPARTMENT OF SCIENCE EDUCATION
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Bachelor of Education in Science **MODULE HANDBOOK**

Module name:	Labwork of General Chemistry 2
Module level, if applicable:	Undergraduate
Code:	IPA6115
Sub-heading, if applicable:	-
Classes, if applicable:	-
Semester:	2th (second)
Module coordinator:	Purwanti Widhy H, M.Pd
Lecturer(s):	Purwanti Widhy H, M.Pd, Putri Anjarsari, M.Pd
Language:	Bahasa Indonesia
Classification within the curriculum:	Compulsory Course
Teaching format / class hours per week during the semester:	100 minutes lectures and 120 minutes structured activities per week.
Workload:	Total workload is 90.67 hours per semester which consists of 100 minutes lectures and 120 minutes structured activities, and 120 minutes individual study per week for 16 weeks.
Credit points:	1 (3 ETCS)
Prerequisites course(s):	
Targeted learning outcomes:	After taking this course the students have ability to: CO1. Show independence and responsible in carrying out individual tasks and group assignments CO2. show independent, systematic and measurable performance CO3. make decisions about solving problems related to chemical experiments consists of identifying groups of organic compounds; test against several chemical compounds such as carbohydrate, fat, protein tests; analysis of electrical conductivity; redox

<table>
<tr><td></td><td>reactions and metal activity series; gilding (electrochemistry), and colloid
CO4. responsible for achieving the results of group work</td></tr>
<tr><td>Content:</td><td>This course contains solving problems related to chemical experiments consists of identifying groups of organic compounds; test against several chemical compounds such as carbohydrate, fat, protein tests; analysis of electrical conductivity; redox reactions and metal activity series; gilding (electrochemistry) and colloid
.</td></tr>
<tr><td>Study / exam achievements:</td><td>Attitude assessment is carried out at each meeting by observation and / or self-assessment techniques using the assumption that basically every student has a good attitude. The student is given a value of very good or not good attitude if they show it significantly compared to other students in general. The result of attitude assessment is not a component of the final grades, but as one of the requirements to pass the course. Students will pass from this course if at least have a good attitude.

The final mark will be weight as follow:

<table>
<tr><th>No</th><th>CO</th><th>Assessment Object</th><th>Assessment Technique</th><th>Weight</th></tr>
<tr><td>1</td><td>CO2, CO3 and CO4</td><td>a. performance
b. pretest
c. report
d. post test</td><td>Presentation / written test</td><td>30%
15%
15%
25%</td></tr>
<tr><td colspan="4">Total</td><td>100%</td></tr>
</table>
</td></tr>
<tr><td>Forms of media:</td><td>Board, LCD Projector, Laptop/Computer</td></tr>
<tr><td>Literature:</td><td>A. Brown, Theodore, et .al, 1976, Chemistry the central science.Pearson: Pearson Pertice Hall.
B. Chang, R., 2004, KIMIA DASAR (konsep-konsep inti), edisi ketiga, jilid 2, Erlangga, Jakarta
C. Keenan, 1989, Kimia untuk Universitas, edisi keenam, jilid 2, Erlangga, Jakarta
D. Silberberg, Martin S. 2006. Principles of General Chemistry. McGraw-Hill Higher Education.</td></tr>
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PLO and CO mapping

	PLO1	PLO2	PLO3	PLO4	PLO5	PLO6	PLO7	PLO8	PLO9	PLO10	PLO11	PLO12
CO1		✓										
CO2								✓				
CO3										✓		
												✓